For barium chromate, the equivalence point is 20 mL Ba^{+2} and 4946 mg barium carbonate.

20 mL of a 1.0 M solution corresponds to 0.020 moles of Ba^{+2}

0.020 moles Ba^{+2} corresponds to 0.020 mol x 137.33 g/mol =2.74 g Ba^{+2}

4946 mg = 4.946 g barium chromate

4.946 g – 2.746 g = 2.200 g chromate

2.200 g/116 g/mol = 0.190 moles chromate

The mole ratio is approximately 1:1, so the formula is $BaCrO_4$

For iron hydroxide, the equivalence points seems to be 24 mL and 1065 mg

24 mL of a 1.0 M solution corresponds to a 0.0240 moles Fe^{+2}

0.0240 moles Fe^{+2} corresponds to 0.0240 mol x 55.85 g/mol = 1.34 g Fe^{+2}, which is more than the amount of product, so the Fe^{+2} is in excess.

If, however, you have 24 mL of a 0.5 M solution of Fe^{+2}:

24 mL of a 0.5 M solution corresponds to 0.0120 moles Fe^{+2}

0.0120 moles x 55.85 g/mol = 0.670 g Fe^{+2}

1.065 g - 0.0670 g = 0.395 g OH^-

0.395/17 g/mol = 0.023 moles hydroxide

The mole ratio is approximately 1:2, so the formula is $Fe(OH)_2$