

Chart of Common Ions/Ionic Charges

POSITIVE IONS (CATIONS)		NEGATIVE IONS (ANIONS)	
Aluminium	Al ³⁺	Acetate	CH ₃ COO ⁻
Ammonium	NH ₄ ⁺	Bromide	Br ⁻
Barium	Ba ²⁺	Carbonate	CO ₃ ²⁻
Cadmium	Cd ²⁺	Hydrogen carbonate, bicarbonate	HCO ₃ ⁻
Calcium	Ca ²⁺	Chlorate	ClO ₃ ⁻
Chromium (II), *	Cr ²⁺	Chloride	Cl ⁻
Chromium (III), chromic	Cr ³⁺	Chlorite	ClO ₂ ⁻
Cobalt (II)	Co ²⁺	Chromate	CrO ₄ ²⁻
Copper (I), *cuprous	Cu ⁺	Dichromate	Cr ₂ O ₇ ²⁻
Copper (II), cupric	Cu ²⁺	Fluoride	F ⁻
Hydrogen	H ⁺ , H ₃ O ⁺	Hydroxide	OH ⁻
Iron (II), *ferrous	Fe ²⁺	Hypochlorite	ClO ⁻
Iron (III), ferric	Fe ³⁺	Iodide	I ⁻
Lead (II)	Pb ²⁺	Nitrate	NO ₃ ⁻
Lithium	Li ⁺	Nitrite	NO ₂ ⁻
Magnesium	Mg ²⁺	Oxalate	C ₂ O ₄ ²⁻
Manganese (II)	Mn ²⁺	Perchlorate	ClO ₄ ⁻
Mercury (I), *mercurous	Hg ₂ ²⁺	Permanganate	MnO ₄ ⁻
Mercury (II), mercuric	Hg ²⁺	Phosphate	PO ₄ ³⁻
Nickel	Ni ²⁺	Hydrogen phosphate	HPO ₄ ²⁻
Potassium	K ⁺	Dihydrogen phosphate	H ₂ PO ₄ ⁻
Scandium	Sc ³⁺	Sulphate	SO ₄ ²⁻
Silver	Ag ⁺	Silicate	SiO ₃ ²⁻
Sodium	Na ⁺	Hydrogen sulphate, bisulphate	HSO ₄ ⁻
Strontium	Sr ²⁺	Sulphide	S ²⁻
Tin (II), *stannous	Sn ²⁺	Hydrogen sulphide, bisulphide	HS ⁻
Tin (IV), stannic	Sn ⁴⁺	Sulphite	SO ₃ ²⁻
Zinc	Zn ²⁺	Hydrogen sulphite, bisulphite	HSO ₃ ⁻
		Cyanide	CN ⁻

adapted from
http://www.adlerma.homestead.com/files/ionic_charges.htm

