Title: Chemical Names and Formulas

Purpose:

- To practice writing chemical names and formulas of common chemical compounds
- To describe the colors and textures of common ionic compounds
- To synthesize chemical compounds

Materials:

- Chemical compound samples
- Prepared solutions of ionic compounds
- Pipettes
- Reaction surface
- Safety glasses

Procedure:

1. Observe and describe the samples of solid chemical compounds. Record your results in a data table.
2. Place 1 drop of each ionic solution on the reaction surface as indicated in the reaction chart.
   Describe the reaction (color change, precipitation, color of precipitate, etc)
   Write the formula for the compound that forms

Results:

Name	Formula	Description
potassium iodide		
sodium chloride		
magnesium sulfate		
copper ii sulfate		
	$NaHCO_3$	
	$AgNO_3$	
	$NaNO_2$	
	KF	
sodium phosphate		
calcium hydroxide		
tin iv chloride		
potassium bromide		
	$CaCl_2$	
	$FeCl_2$	
	Na_2HPO_4	
	NaH_2PO_4	

	$AgNO_3$ (Ag^+) (toxic)	$Pb(NO_3)_2$ (Pb^{+2}) (toxic)	$CuSO_4$ (Cu^{+2})	$MgSO_4$ (Mg^{+2})	$FeCl_3$ (Fe^{+3})
$FeCl_3$ (Cl^-)					
KI (I^-)					
NaOH (OH^-) (toxic)					
Na_2CO_3 (CO_3^{-2})					
Na_3PO_4 (PO_4^{-3})					

	$AgNO_3$ (Ag^{+}) (toxic)	$Pb(NO_3)_2$ (Pb^{+2}) (toxic)	$CuSO_4$ (Cu^{+2})	$MgSO_4$ (Mg^{+2})	$FeCl_3$ (Fe^{+3})
$FeCl_3$ (Cl^{-})	X	X			
KI (I^{-})	X	X			
NaOH (OH^{-}) (toxic)	X	X	X	X	X
Na_2CO_3 (CO_3^{-2})	X	X	X	X	X
Na_3PO_4 (PO_4^{-3})	X	X	X	X	X