Chemistry: Moles and Avogadro's Number

To determine mass of more than 1 mole, multiply formula mass (molar mass) by number of moles.
To determine number of particles, multiply number of moles by Avogadro's number (6.02×10^{23})
    Particles may be atoms, molecules, formula units, etc.

Compound	Molar Mass (g)	Particles 1.00 mole	Mass 2.00 moles (g)	Particles 2.00 moles	Mass 0.50 Moles (g)	Particles 0.50 moles	Mass 1.50 moles (g)	Particles 1.50 moles
KCl								
KOH								
$KMnO_4$								
KNO_3								
K_2SO_4								
$AgCl$								
SnF_2								
NH_4OH								
Na_2CO_3								
CaO								
$Ca(OH)_2$								
$AgNO_3$								
CO_2								
CCl_4								
NH_3								
CH_4								
O_2								
H_2								
Cl_2								
I_2								